

# Chemistry 2720 Fall 2003 Quiz 4

Name: \_\_\_\_\_

Calculate the partial pressure of free chlorine atoms in gaseous chlorine at 25°C and a total pressure of 1.3 bar. The standard free energy of formation of a chlorine atom in the gas phase is 105.305 kJ/mol.

To convert degrees Celsius to Kelvin, add 273.15.

$$R = 8.314472 \text{ J K}^{-1} \text{ mol}^{-1}$$